

Example Of One Step Problems - Mass

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How many grams of carbon (C) do I have 4.35 moles of C ?

1. ? g C

2. 4.35 mole C

$$\frac{12 \text{ g C}}{1 \text{ mole C}}$$

called atomic mass

Equation: #mole =

$$\frac{\# \text{ g}}{\left(\frac{\# \text{ g}}{\text{mole}}\right)}$$

3.

DA:

$$= (4.35 \text{ mole C}) \left(\frac{12 \text{ g C}}{1 \text{ mole C}} \right)$$

$$= 52.20 \text{ g C}$$

$$= 52.2 \text{ g C}$$

$$\# \text{ g} = (\# \text{ mole}) \left(\frac{\# \text{ g}}{\text{mole}} \right)$$

plug in and get this

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