

Example Of One Step Mole Problems - Numbers

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How many atoms of He do I have if I have 3.56 moles of He?

1. ? atom He

2. 3.56 mole He

$$\frac{6.02 \times 10^{23} \text{ atom He}}{1 \text{ mole He}}$$

DA:

$$x = (3.56 \text{ mole He}) \left(\frac{6.02 \times 10^{23} \text{ atom He}}{1 \text{ mole He}} \right)$$

$$x = 2.143 \times 10^{24} \text{ atom He}$$

$$x = 2.14 \times 10^{24} \text{ atom He}$$

Equation.

$$\# \text{ mole} = \frac{\# \text{ atom}}{\left(\frac{6.02 \times 10^{23} \text{ atom}}{1 \text{ mole}} \right)}$$

$$\# \text{ atom} = (\# \text{ mole}) \left(\frac{6.02 \times 10^{23} \text{ atom}}{1 \text{ mole}} \right)$$

Rest of answer is
same as DA.